

Australian Standard<sup>®</sup>

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**WATERS—  
DETERMINATION OF FILTRABLE  
SYNTHETIC ANIONIC  
SURFACTANTS—  
COPPER-ETHYLENEDIAMINE  
FLAME ATOMIC ABSORPTION  
SPECTROMETRIC METHOD**

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This Australian Standard was prepared by Committee CH/22, Methods for Examination of Water. It was approved on behalf of the Council of the Standards Association of Australia on 2 September 1987 and published on 2 November 1987.

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The following interests are represented on Committee CH/22:

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Confederation of Australian Industry  
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Department of Housing and Construction  
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*This Standard was issued in draft form for comment as DR 85023.*

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First published as AS 3506 . . . . . 1987

PUBLISHED BY STANDARDS AUSTRALIA  
(STANDARDS ASSOCIATION OF AUSTRALIA)  
1 THE CRESCENT, HOMEBUSH, NSW 2140

ISBN 0 7262 4755 3

## PREFACE

This Standard for the determination of filtrable synthetic anionic surfactants in waters was prepared by the Association's Committee on Methods for Examination of Waters under the direction of the Chemical Standards Board.

Synthetic anionic surfactants are determined in waters because they may be present in concentrations which are harmful to aquatic biota or produce undesirable foaming. This method is simpler than the methylene blue method and suffers fewer interferences (Appendix A). The response to the method of linear alkylsulphates of different chain lengths correlates more closely with the foaming characteristics of these compounds than does their response to either the methylene blue or azure A methods (Appendix B).

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## STANDARDS ASSOCIATION OF AUSTRALIA

## Australian Standard

WATERS—DETERMINATION OF FILTRABLE SYNTHETIC ANIONIC SURFACTANTS—  
COPPER-ETHYLENEDIAMINE FLAME ATOMIC ABSORPTION  
SPECTROMETRIC METHOD

**1 SCOPE.** This Standard sets out a flame atomic absorption spectrometric method for the determination of filtrable synthetic anionic surfactants in waters.

**2 APPLICATION.** The method is applicable to natural and waste waters that contain synthetic anionic surfactants in the concentration range 0.02 mg/L to 0.5 mg/L (expressed as linear alkylbenzene sulphonate, (LAS)). The lower limit of the method depends on instrument sensitivity (see Note to Clause 9.2(1)) and precision (Clause 12). The upper limit may be extended by taking smaller test sample volumes.

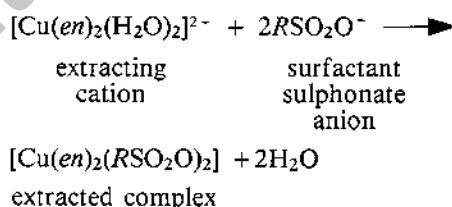
NOTE: Information on interferences by common cations and anions is given in Appendix A.

**3 REFERENCED DOCUMENTS.** The following Standards are referred to in this Standard:

- AS 2031.1 Selection of Containers and Preservation of Water Samples for Chemical and Microbiological Analysis.  
Part 1: Chemical.
- AS 2134.1 Recommended Practice for Chemical Analysis by Atomic Absorption Spectrometry  
Part 1: Flame Atomic Absorption Spectrometry.
- AS 2162 Code of Practice for the Use of Volumetric Glassware.
- AS 2407 Separating Funnels and Dropping Funnels.
- AS 2850 Chemical Analysis—Interlaboratory Test Programs—Guide to Planning and Conduct—For Determining Precision of Analytical Methods(s).

**4 PRINCIPLE.** At pH 5–9, neutral complexes form between synthetic surfactant anions and the bis(ethylenediamine)copper(II) ion. These complexes are extracted into chloroform and the copper they contain is back-extracted into dilute acid. The copper in the acid extract is determined by flame atomic absorption spectrometry.

**5 REACTION.** Formation of the extractable complex may be represented by the following example:



where

*en* = ethylenediamine ( $\text{H}_2\text{NCH}_2\text{CH}_2\text{NH}_2$ )

*R* = hydrocarbon group.

NOTE: Further information on the nature of the extracting agent and the extracted species is given in Appendix C.

## 6 REAGENTS.

**6.1 General requirements.** Unless otherwise specified, use analytical grade reagents and distilled water or water of equivalent purity.

### 6.2 Solutions.

#### 6.2.1 Chloroform.

**SAFETY WARNING.** Care should be taken to avoid breathing chloroform vapours during the handling operations in a fume cupboard) and to avoid contact with the skin.

**6.2.2 Copper-ethylenediamine reagent.** Dissolve 62 g of copper sulphate pentahydrate ( $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ ), 50 g of ammonium sulphate ( $(\text{NH}_4)_2\text{SO}_4$ ) and 45 g of ethylenediamine in water and dilute to 1 L.

**6.2.3 Hydrochloric acid (approx. 0.1 mol/L).** Dilute 10 mL of hydrochloric acid ( $\rho_{20}$  1.16 g/mL) to 1 L.

**6.2.4 Nitric acid (1 + 4).** Add 200 mL of nitric acid ( $\rho_{20}$  1.42 g/mL) to 800 mL of water.

### 6.3 Standard solutions.

**6.3.1 Stock standard LAS solution (approx. 1000 mg/L).** Weigh, to the nearest 0.1 mg, sufficient LAS standard to give on dilution, a solution containing approximately 1000 mg LAS/L. Calculate the concentration of the stock standard solution to four significant figures from the following formula:

$$c_{\text{stock(mg/L)}} = \frac{c \times m \times M \times 1000}{V_1}$$

where

*c* = concentration of LAS in the LAS standard, in millimoles per gram (mmol/g)

*m* = mass of LAS standard weighed, in grams

*M* = molecular weight of the LAS, as the sulphonic acid

*V*<sub>1</sub> = volume to which the LAS is diluted, in millilitres.

NOTES:

- Enquiries regarding the supply of LAS standard may be directed to the Curator of Standards, Australian Government Analytical Laboratory, 11 William Street, Melbourne, Victoria 3000. LAS standard is supplied as a solution of the triethanolamine salt of Dobane 124AS sulphonate (molecular weight of the LAS, as the sulphonic acid = 323).
- The stock standard LAS solution is stable for one week when stored at 4°C in the dark.